

Sri Sathya Sai College for Women, Bhopal

(An Autonomous College affiliated to Barkatullah University, Bhopal)

(NAAC Accredited 'A' Grade)



SYLLABUS

PG

SESSION- 2023-24

Class: M.Sc. Semester-I & II

SUBJECT: Chemistry

Dr. Vashu
Vashu
Dr. Vashu
Anita Dutt Karan
Anita
Vashu

Sri Sathya Sai College for Women, Bhopal
(An Autonomous College Affiliated to Barkatullah University Bhopal)

Department of Higher Education, Govt. of M.P.

Post Graduate Semester Wise Syllabus

As recommended by Central Board of Studies and approved by the Governor of M. P.
(2023-24)

Max Marks : 100 (Theory - 85 + CCE- 15)
Class/Semester : M.Sc. I
Subject : Chemistry
Paper : I (Compulsory)
Title : INORGANIC CHEMISTRY- I

Unit-I Stereochemistry and Bonding in Main Group Compounds :

VSEPR, Walsh diagram (triatomic and penta-atomic molecules), $d\pi-p\pi$ bond, Bent rule and energetics of hybridization, some simple reactions of covalently bonded molecules.

Unit-II Metal-Ligand Equilibrium in Solution

Stepwise and overall formation constants and their interaction, trends in stepwise constant, factors affecting the stability of metal complexes with reference to the nature of metal ion and ligand. Chelate effect and its thermodynamic origin, determination of binary formation constants by potentiometry and spectrophotometry.

Unit-III Reaction Mechanism of Transition Metal Complexes

Energy profile of a reaction, reactivity of metal complex, inert and labile complexes, kinetic application of valence bond and crystal field theories, kinetics of octahedral substitution, acid hydrolysis, factors affecting acid hydrolysis, base hydrolysis, conjugate base mechanism, direct and indirect evidences in favour of conjugate mechanism, anion reactions, reactions without metal ligand bond cleavage. Substitution reactions in square planar complexes, the trans effect, mechanism of the substitution reaction. Redox reaction, electron transfer reactions, mechanism of one electron transfer reactions, outer sphere type reactions, cross reactions and Marcus-Hush theory, inner sphere type reactions.

Unit-IV Metal-Ligand bonding

Limitation of crystal field theory, molecular orbital theory for bonding in octahedral, tetrahedral and square planar complexes, π -bonding and molecular orbital theory.

Unit-V- HSAB Theory : Classification of acids and bases as hard and soft; HSAB principle, theoretical basis of hardness and softness; Lewis-acid base reactivity approximation; donor and acceptor numbers, E and C equation; applications of HSAB concept.

Handwritten signatures and notes:
A large signature on the left.
A signature in the middle with the word "heno" written above it.
A signature on the right.
A large signature at the bottom right that reads "Anita Dutt Kour".
A small signature at the bottom right corner that reads "Vashta".

Scheme of Question Paper

Type of Question	No. of Question	Marks Alloted	Total
Objective	10	1	10
Short Answer	5	5	25
Long Answer	5	10	50
Total			85

Books

Suggested :

1. Advanced Inorganic Chemistry, F.A. Cotton and Wilkinson, John Wiley.
2. Inorganic Chemistry, J.E. Huhey, Harpes & Row.
3. Chemistry of the Elements. N.N. Greenwood and A. Earnshaw, Pergamon.
4. Inorganic Electronic Spectroscopy, A.B.P. Lever, Elsevier.
5. Magnetochemistry, R.I. Carlin, Springer Verlag.
6. Comprehensive Coordination Chemistry eds., G. Wilkinson, R.D. Gillars and J.A. Mc Cleverty, Pergamon.

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Max Marks	: 100 (Theory- 85+ CCE- 15)
Class/Semester	: M.Sc. I
Subject	: Chemistry
Paper	: II (Compulsory)
Title	: ORGANIC CHEMISTRY I

Unit-I

Nature of Bonding in Organic Molecules

Delocalized chemical bonding-conjugation, cross conjugation, resonance, hyperconjugation, bonding in fullerenes, tautomerism. Aromaticity in benzenoid and non-benzenoid compounds, alternate and non-alternate hydrocarbons. Huckel's rule, energy level of π -molecular orbitals, annulenes, anti-aromaticity, homo-aromaticity, PMO approach. Bonds weaker than covalent-addition compounds, crown ether complexes and cryptands, inclusion compounds, catenanes and rotaxanes.

Unit-II

Stereochemistry

Strain due to unavoidable crowding Elements of symmetry, chirality, molecules with more than one chiral center, threo and erythro isomers, methods of resolution, optical purity, enantiotopic and diastereotopic atoms, groups and faces, stereospecific and stereoselective synthesis, Asymmetric synthesis. Optical activity in the absence of chiral carbon (biphenyls, allenes and spirane) chirality due to helical shape. Stereochemistry of the compounds containing nitrogen, sulphur and phosphorus.

Unit III

Conformational analysis and linear free energy relationship

Conformational analysis of cycloalkanes, decalines, effect of conformation on reactivity, conformation of sugars. Generation, structure, stability and reactivity of carbocations, carbanions, free radicals, carbenes and nitrenes. The Hammett equation and linear free energy relationship, substituents and reaction constants, Taft equation.

Unit-IV

Reaction Mechanism : Structure and Reactivity

Type of mechanisms, types of reactions, thermodynamic and kinetic requirements, kinetic and thermodynamic control, Hammond's postulate, Curtin-Hammett principle. Potential energy diagrams, transition states and intermediates, methods of determining mechanisms, isotope effects

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Unit-V

Aliphatic Nucleophilic Substitution

The S_N2 , S_N1 mixed S_N1 and S_N2 and SET mechanism. The neighboring group mechanism, neighboring group participation by σ and π bonds, anchimeric assistance. Classical and nonclassical carbocations, phenonium ions, norbornyl systems, common carbocation rearrangements. Application of NMR spectroscopy in the detection of carbocations. The S_N1 mechanism. Nucleophilic substitution at an allylic, aliphatic trigonal and a vinylic carbon. Reactivity effects of substrate structure, attacking nucleophile, leaving group and reaction medium, phase transfer catalysis and ultrasound, ambident nucleophile, regioselectivity.

Scheme of Question Paper

Type of Question	No. of Question	Marks Alloted	Total marks
Objective	10	1	10
Short Answer	5	5	25
Long Answer	5	10	50
Total			85

Book Suggested

1. Advanced Organic Chemistry-Reactions, Mechanism and Structure, Jerry March, John Wiley.
2. Advanced Organic Chemistry, F.A. Carey and R.J. Sundberg, Plenum.
3. A Guide Book to Mechanism in Organic Chemistry, Peter Sykes, Longman.
4. Structure and Mechanism in Organic Chemistry, C.K. Ingold, Cornell University Press.
5. Organic Chemistry, R.T. Morrison and R.N. Boyd, Prentice-Hall.
6. Modern Organic Reactions, H.O. House, Benjamin.
7. Principles of Organic Synthesis, R.O.C. Norman and J.M. Coxon, Blackie Academic & Professionals.
8. Reaction Mechanism in Organic Chemistry, S.M. Mukherji and S.P. Singh, Macmillan.
9. Pericyclic Reactions, S.M. Mukherji, Macmillan, India
10. Stereochemistry of Organic Compounds, D.Nasipuri, New Age International.
11. Stereochemistry of Organic Compounds, P.S. Kalsi, New Age International.

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Max Marks	: 100 (Theory- 85+ CCE- 15)
Class/Semester	: M.Sc. I
Subject	: Chemistry
Paper	: III (Compulsory)
Title	: PHYSICAL CHEMISTRY - I

Unit-I

Introduction to Exact Quantum Mechanical Results

Schrödinger equation and the postulates of quantum mechanics. Discussion of solutions of the Schrödinger equation to some model systems viz., particle in a box, the harmonic oscillator, the rigid rotor, the hydrogen atom and helium atom.

Unit-II

Approximate Methods

The variation theorem, linear variation principle, Perturbation theory (First order and nondegenerate). Applications of variation method and perturbation theory to the Helium atom.

Molecular Orbital Theory

Huckel theory of conjugated systems bond and charge density calculations. Applications to ethylene, butadiene, cyclopropenyl radical, cyclobutadiene etc. Introduction to extended Huckel theory.

UNIT III

Angular Momentum

Ordinary angular momentum, generalized angular momentum, eigenfunctions for angular momentum, eigenvalues of angular momentum operator using ladder operators, addition of angular momenta, spin, antisymmetry and Pauli exclusion principle.

Unit-IV

Classical Thermodynamics

Brief resume of concepts of laws of thermodynamics, free energy, chemical potential and entropies. Partial molar free energy, partial molar volume and partial molar heat content and their significance. Determinations of these quantities. Concept of fugacity and determination of fugacity. Non-ideal systems : Excess functions for non-ideal solutions. Activity, activity coefficient, Debye Huckel theory for activity coefficient of electrolytic solutions; determination of activity and activity coefficients; ionic strength. Application of phase rule to three component systems; second order phase transitions.

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Unit-V

Statistical Thermodynamics

Concept of distribution, thermodynamic probability and most probable distribution. Ensemble averaging, postulates of ensemble averaging. Canonical, grand canonical and micro-canonical ensembles, corresponding distribution laws (using Lagrange's method of undetermined multipliers). Partition functions-translation, rotational, vibrational and electronic partition functions. Calculation of thermodynamic properties in terms of partition. Application of partition functions. Fermi-Dirac Statistics, distribution law and applications to metal. Bose-Einstein statistics distribution Law and application to helium.

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1. Physical Chemistry, P.W. Atkins, ELBS.
2. Introduction to Quantum Chemistry, A.K. Chandra, Tata Mc Graw Hill.
3. Quantum Chemistry, Ira N. Levine, Prentice Hall.
4. Coulson's Valence, R. Mc Ween y, ELBS.
5. Chemical Kinetics. K.J. Laidler, McGraw-Hill.
6. Kinetics and Mechanism of Chemical Transformation J.Rajaraman and J. Kuriacose, Mc Millan.
7. Micelles, Theoretical and Applied Aspects, V. MOraoi, Plenum.
8. Modern Electrochemistry Vol. I and Vol II J.O.M. Bockris and A.K.N. Reddy, Planum.
9. Introduction to Polymer Science, V.R. Gowarikar, N.V. Vishwanathan and J. Sridhar, Wiley Eastern.
10. Introduction to Quantum Chemistry-R.K. Prasad, New Age Publication.

Answer -

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Max Marks	: 100 (Theory- 85+ CCE- 15)
Class/Semester	: M.Sc. I
Subject	: Chemistry
Paper	: IV (Compulsory)
Title	: Group Theory & Spectroscopy - I

Unit-I - Symmetry and Group theory in Chemistry

Symmetry elements and symmetry operation, definition of group, subgroup. Conjugacy relation and classes. Point symmetry group. Schoenflies symbols, representations of groups by matrices (representation for the C_n , C_{nv} , C_{nh} , D_{nh} group to be worked out explicitly). Character of a representation. The great orthogonality theorem (without proof) and its importance. Character tables and their use; spectroscopy. Derivation of character table for C_{2v} and C_{3v} point group. Symmetry aspects of molecular vibrations of H_2O molecule.

Unit-II - Microwave Spectroscopy

Classification of molecules, rigid rotor model, effect of isotopic substitution on the transition frequencies, intensities, non-rigid rotor. Stark effect, nuclear and electron spin interaction and effect of external field. applications.

Unit-III - Infrared-Spectroscopy

Review of linear harmonic oscillator, vibrational energies of diatomic molecules, zero point energy, force constant and bond strengths; anharmonicity, Morse potential energy diagram, vibration-rotation spectroscopy. P.Q.R. branches, Breakdown of Oppenheimer approximation; vibrations of polyatomic molecules. Selection rules, normal modes of vibration, group frequencies, overtones, hot bands, factors affecting the band positions and intensities, far IR region, metal ligand vibrations, normal co-ordinate analysis.

Unit-IV - Raman Spectroscopy

Classical and quantum theories of Raman effect. Pure rotational, vibrational and vibrational-rotational Raman spectra, selection rules, mutual exclusion principle, Resonance Raman spectroscopy, coherent anti stokes Raman spectroscopy (CARS).

Unit-V- Electronic Spectroscopy

Molecular Spectroscopy

Energy levels, molecular orbitals, vibronic transitions, vibrational progressions and geometry of the excited states, Franck-Condon principle, electronic spectra of polyatomic molecules. Emission spectra; radio-active and non-radioactive decay, internal conversion, spectra of transition metal complexes, charge-transfer spectra.

Photoelectron Spectroscopy Basic principles; photo-electric effect, ionization process, Koopman's theorem. Photoelectron spectra of simple molecules, ESCA, chemical information from ESCA. Auger electron spectroscopy-basic idea.

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Objective	10	1	10
Short Answer	5	5	25
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Total			85

Books suggested

1. Modern Spectroscopy, J.M. Hollas, John Wiley.
2. Applied Electron Spectroscopy for chemical analysis d. H. Windawi and F.L. Ho, Wiley Interscience.
3. NMR, NQR, EPr and Mossbauer Spectroscopy in Inorganic Chemistry, R.V.Parish, Ellis Harwood.
4. Physical Methods in Chemistry, R.S. Drago, Saunders College.
5. Chemical Applications of Group Theory, F.A. Cotton.
6. Introduction to Molecular Spectroscopy, G.M. Barrow, Mc Graw Hill.
7. Basic Principles of Spectroscopy, R. Chang, Mc Graw Hill.
8. Theory and Application of UV Spectroscopy, H.H. Jaffe and M. Orchin, IBHOxford.
9. Introduction to Photoelectron Spectroscopy, P.K. Ghosh, John Wiley.
10. Introduction to Magnetic Resonance. A Carrington and A.D. Maclachalan, Harper & Row.

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Post Graduate Semester Wise Syllabus

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Max Marks : 100 (Theory- 85+ CCE- 15)
Class/Semester : M.Sc. I
Subject : Chemistry
Paper : V (Compulsory)
Title : MATHEMATICS FOR CHEMISTS

(For students without Mathematics in B.Sc.)

Unit-I- Vectors

Vectors, dot, cross and triple products etc. gradient, divergence and curl, Vector Calculus.

Matrix Algebra

Addition and multiplication; inverse, adjoint and transpose of matrices.

Unit-II -bDifferential Calculus

Functions, continuity and differentiability, rules for differentiation, applications of differential calculus including maxima and minima (examples related to maximally populated rotational energy levels, Bohr's radius and most probable velocity from Maxwell's distribution etc.).

Unit-II - Integral calculus

Basic rules for integration, integration by parts, partial fractions and substitution. Reduction formulae, applications of integral calculus.

Functions of several variables, partial differentiation, co-ordinate transformations (e.g. Cartesian to spherical polar).

Unit-IV - Elementary Differential equations

First-order and first degree differential equations, homogenous, exact and linear equations. Applications to chemical kinetics, secular equilibria, quantum chemistry etc. second order differential equation and their solutions.

Unit-v Permutation and Probability

Permutations and combinations, probability and probability theorems average, variance root means square deviation examples from the kinetic theory of gases etc., fitting(including least squares fit etc with a general polynomial fit.

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Objective	10	1	10
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Long Answer	5	10	50
Total			85

Book Suggested

1. The chemistry Mathematics Book, E.Steiner, Oxford University Press.
2. Mathematics for chemistry, Doggett and Suiclific, Logman.
3. Mathematical for Physical chemistry : F. Daniels, Mc. Graw Hill.
4. Chemical Mathematics D.M. Hirst, Longman.
5. Applied Mathematics for Physical Chemistry, J.R. Barante, Prentice Hall.
6. Basic Mathematics for Chemists, Tebbutt, Wiley.

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Max Marks : 100 (Theory 85+ CCE 15)
Class/Semester : M.Sc. I
Subject : Chemistry
Paper : V (Compulsory)
Title : BIOLOGY FOR CHEMISTS

(For students without Biology in B.Sc.)

Unit-I- Cell Structure and Functions

Structure prokaryotic and eukaryotic cells, intracellular organelles and their functions, comparison of plant and animal cells. Overview and their functions, comparison of plant and animal cells. Overview of metabolic processes-catabolism and anabolism. ATP – the biological energy currency. Origin of life-unique properties of carbon chemical evolution and rise of living systems. Introduction to bio-molecules, building blocks of biomacromolecules.

Unit-II- Carbohydrates

Conformation of monosaccharides, structure and functions of important derivatives of monosaccharides like glycosides, deoxy sugars, myoinositol, amino sugars. N-acetylmuramic acid, sialic acid disaccharides and polysaccharides. Structural polysaccharides cellulose and chitin. Storage polysaccharides-starch and glycogen. Structure and biological function of glucosaminoglycans of mucopolysaccharides. Carbohydrates of glycoproteins and glycolipids. Role of sugars in biological recognition. Blood group substances. Ascorbic acid.

Unit-III - Lipid

Fatty acids, essential fatty acids, structure and function of triacylglycerols, glycerophospholipids, sphingolipids, cholesterol, bile acids, prostaglandins. Lipoproteins-composition and function, role in atherosclerosis. Properties of lipid aggregates-micelles, bilayers, liposomes and their possible biological functions. Biological membranes. Fluid mosaic model of membrane structure. Lipid metabolism-oxidation of fatty acids.

Unit-IV- Amino-acids, Peptides and Proteins

Chemical and enzymatic hydrolysis of proteins to peptides, amino acid sequencing. Secondary structure of proteins. force responsible for holding of secondary structures. α -helix, β -sheets, super secondary structure, triple helix structure of collagen. Tertiary structure of protein-folding and domain structure. Quaternary structure. Amino acid metabolism-degradation and biosynthesis of amino acids, sequence determination : chemical/enzymatic/mass spectral, racemization/detection. Chemistry of oxytocin and tryptophan releasing hormone (TRH).

Anita Sult Kaur
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Unit-V - Nucleic Acids

Purine and pyrimidine bases of nucleic acids, base pairing via H bonding. Structure of ribonucleic acids (RNA) and deoxyribonucleic acid (DNA), double helix model of DNA and forces responsible for holding it. Chemical and enzymatic hydrolysis of nucleic acids. The chemical basis for heredity, an overview of replication of DNA, transcription, translation and genetic code. Chemical synthesis of mono and trinucleoside.

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Objective	10	1	10
Short Answer	5	5	25
Long Answer	5	10	50
Total			85

Book Suggested

1. Principles of Biochemistry, A.L. Lehninger, Worth Publishers.
2. Biochemistry, L. Stryer, W.H. Freeman.
3. Biochemistry, J. David Rawan, Neil Patterson.
4. Biochemistry, Voet and Voet, John Wiley.
5. Outlines of Biochemistry E.E. Conn and P.K. Stumpf, John Wiley.

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Post Graduate Semester Wise Syllabus

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(2023-24)

Class & Semester : M.Sc. I
Subject : Chemistry
Paper : Chemistry Practical
Title : Inorganic Chemistry

Max Marks : 33

Time: 6-8hours

Qualitative analysis	8
Chromatography	8
Preparation	8
Record	4
Viva Voce	5

1. Qualitative Analysis

Qualitative analysis of inorganic mixture containing eight radicals including insoluble and rare earth.

2. Chromatography

Separation of cations and anions by Column Chromatography: Ion exchange

Separation of cations and anions by Paper Chromatography.

3. Preparations

Preparation of selected inorganic compounds and their studies by I.R. electronic spectra, Mossbauer, E.S.R. and magnetic susceptibility measurements. Handling of air and moisture sensitive compounds.

1. $\text{VO}(\text{acac})_2$
2. $\text{TiO}(\text{C}_9\text{H}_8\text{NO})_2\text{H}_2\text{O}$
3. $\text{cis-K}[\text{Cr}(\text{C}_2\text{O}_4)_2(\text{H}_2\text{O})_2]$
4. $\text{Na}[\text{Cr}(\text{NH}_3)_2(\text{SCN})_4]$
5. $\text{Ni}(\text{acac})_2$
6. $\text{K}_3[\text{Fe}(\text{C}_2\text{O}_4)_3]$
7. Prussian Blue, Turnbull's Blue

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Anita Dutt Kaur
Vishu Datta
K. S.
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Post Graduate Semester Wise Syllabus

As recommended by Central Board of Studies and approved by the Governor of M. P.
(2023-24)

Class & Semester : M.Sc. I
Subject : Chemistry
Paper : Chemistry Practical
Title : Organic Chemistry

Max Marks : 33

Time: 6-8hours

Qualitative Analysis 12
Organic Synthesis 12
Record 4
Viva Voce 5

Qualitative Analysis

Separation, purification and identification of compounds of binary mixture (one liquid and one solid) using TLC and columns chromatography, chemical tests. IR spectra to be used for functional group identification.

Organic Synthesis

Synthesis of organic compounds by applying following methods

1. Acetylation :
2. Oxidation
3. Halogenation
4. Benzoylation
5. Reduction

The Products may be Characterized by Spectral Techniques.

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Class & Semester : M.Sc. I
Subject : Chemistry
Paper : Chemistry Practical
Title : Physical Chemistry

Max Marks : 34

Time: 6-8hours

Error Analysis and Statistical Data Analysis	8
Conductometry	9
Solution	8
Record	4
Viva Voce	5

Error Analysis and Statistical Data Analysis

Errors, types of errors, minimization of errors distribution curves precision, accuracy and combination; statistical treatment for error analysis, student's t-test, null hypothesis, rejection criteria. F & Q test; linear regression analysis, curve fitting. Calibration of volumetric apparatus, burette, pipette and standard flask.

Conductometry

- Determination of the velocity constant, order of the reaction and energy of activation for saponification of ethyl acetate by sodium hydroxide conductometrically.
- Determination of solubility and solubility product of sparingly soluble salts (e.g. PbSO_4 , BaSO_4) conductometrically.
- Determination of the strength of strong and weak acid in a given mixture conductometrically.
- To study of the effect of solvent on the conductance of AgNO_3 /acetic acid and to determine the degree of dissociation and equilibrium constant in different solvents and in their mixtures (DMSO, DMF, dioxane, acetone, water) and to test the validity of Debye-Huckel-Onsager theory.
- Determination of the activity coefficient of zinc ions in the solution of 0.002 M zinc sulphate using Debye Huckel's limiting law.

Solution

Determination of molecular weight of non-volatile and electrolyte/electrolyte by cryoscopic method and to determine the activity coefficient of an electrolyte.

Determination of the degree of dissociation of weak electrolyte and to study the deviation from ideal behaviour that occurs with a strong electrolyte.

Adsorption :To study surface tension-concentration relationship for solutions (Gibbs equation)
Phase Equilibria

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- i. Determination of congruent composition and temperature of a binary system (e.g. diphenylamine-benzophenone system).
- ii. To construct the phase diagram for three component system (e.g. chloroform-acetic acid-water).

Books Suggested

1. Vogel's Textbook of Quantitative Analysis, revised, J. Bassett, R.C. Denney, G.H. Jeffery and J. Mendham, ELBS.
2. Synthesis and Characterization of Inorganic Compounds, W.L. Jolly. Prentice Hall.
3. Experiments and Techniques in Organic Chemistry, D.P. Pasto, C. Johnson and M. Miller, Prentice Hall.
4. Macroscale and Microscale Organic Experiments, K.L. Williamson, D.C. Heath.
5. Systematic Qualitative Organic Analysis, H. Middleton, Edward Arnold.
6. Handbook of Organic Analysis-qualitative and Quantitative. H. Clark, Edward Arnold.
7. Vogel's Textbook of Practical Organic Chemistry, A.R. Tatchell, John Wiley.
8. Practical Physical Chemistry, A.M. James and F.E. Prichard, Longman.
9. Findley's Practical Physical chemistry, B.P. Levitt, Longman.
10. Experimental Physical Chemistry, R.C. Das and B. Behera, Tata McGraw Hill.

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Subject : Chemistry
Paper : I (Compulsory)
Title : Inorganic Chemistry

Unit-I

Electronic Spectral Studies of Transition Metal Complexes :

Spectroscopic ground states, correlation. Orgel and Tanabe-Sugano diagrams for transition metal complexes (d^1 - d^9 states), Selection rule for electronic spectroscopy. Intensity of various type electronic transitions. Calculations of $10Dq$, B and β parameters, charge transfer spectra.

Unit-II

Magnetic Properties of Transition Metal Complexes

Anomalous magnetic moments, Quenching of Orbital contribution. Orbital contribution to magnetic moment, magnetic exchange coupling and spin crossover.

Unit-III

Metal π -Complexes

Metal carbonyl, structure and bonding, vibrational spectra of metal carbonyls for bonding and structural elucidation, important reactions of metal carbonyls : preparation, bonding, structure and important reaction of transition metal nitrosyl, dinitrogen and dioxygen complexes, tertiary phosphine as ligand.

Unit-IV

Metal Clusters

Higher boranes, carboranes, metallo-boranes and metallo-carboranes compounds with metal metal multiple bonds.

Unit-V

Optical Rotatory Dispersion and Circular Dichroism

Linearly and circularly polarized lights; optical rotatory power and circular birefringence, ellipticity and circular dichroism; ORD and Cotton effect, Faraday and Kerr effects Assignment of electronic transitions; applications of ORD and CD for the determination of

(i) absolute configuration of complexes and (ii) isomerism due to non-planarity of chelate rings.

Dr.

Dr.

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Dr.

Anita Suthakar

Nasha Khan

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3. Chemistry of the Elements. N.N. Greenwood and A. Earnshaw, Pergamon.
4. Inorganic Electronic Spectroscopy, A.B.P. Lever, Elsevier.
5. Magnetochemistry, R.I. Carlin, Springer Verlag.
6. Comprehensive Coordination Chemistry eds., G. Wilkinson, R.D. Gillars and J.A. Mc Cleverty, Pergamon.

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Handwritten signature: Anita Dutt Kaur

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Handwritten signature: Nisha

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(2023-24)

Max Marks	: 100 (Theory 85+ CCE 15)
Class/Semester	: M.Sc. II
Subject	: Chemistry
Paper	: II (Compulsory)
Title	: Organic Chemistry <u>II</u>

Unit-I- Aromatic Electrophilic Substitution

The arenium ion mechanism, orientation and reactivity, energy profile diagrams. The ortho/para ratio, ipso attack, orientation in other ring systems. Quantitative treatment of reactivity in substrates and electrophiles. Diazonium coupling, Vilsmeier reaction, Gatterman-Koch reaction

Aromatic Nucleophilic Substitution

The S_NAr S_N1 , benzyne and S_N1 mechanism, Reactivity effect of substrate structure, leaving group and attacking nucleophile. The Von Richter, Sommelet-Hauser, and Smiles rearrangements.

Unit-II- Free Radical Reactions

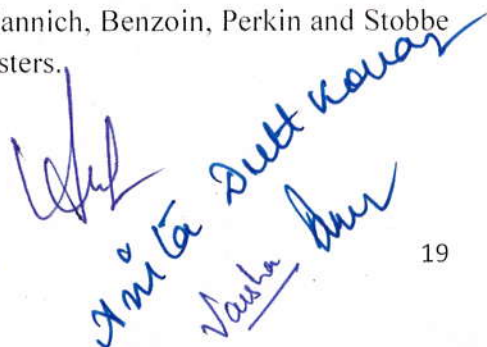
Types of free radical reactions, free radical substitution mechanism, mechanism at an aromatic substrate, neighbouring group assistance. Reactivity for aliphatic and aromatic substrates at a bridgehead. Reactivity in the attacking radicals. The effect of solvents on reactivity. Allylic halogenation (NBS), oxidation of aldehydes to carboxylic acids, autooxidation, coupling of alkynes and arylation of aromatic compounds by diazonium salts, Sandmeyer reaction. Free radical rearrangement. Hunsdiecker reaction.

Unit III- Addition Reactions - Mechanistic and stereochemical aspects of addition reactions involving electrophiles, nucleophiles and free radicals, regio- and chemoselectivity, orientation and reactivity. Addition to cyclopropane ring. Hydrogenation of double and triple bonds, hydrogenation of aromatic rings. Hydroboration, Michael reaction, Sharpless asymmetric epoxidation.

Unit-IV- Addition to Carbon-Hetero Multiple bonds

Mechanism of metal hydride reduction of saturated and unsaturated carbonyl compounds, acid esters and nitriles. Addition of Grignard reagents, organozinc and organolithium reagents to carbonyl and unsaturated carbonyl compounds. Wittig reaction. Mechanism of condensation reactions involving enolates-Aldol, Knoevenagel, Claisen, Mannich, Benzoin, Perkin and Stobbe reactions. Hydrolysis of esters and amides, ammonolysis of esters.

Elimination Reactions



Amita Dubey
Vandana

The E2, E1 and E1cB mechanisms and their spectrum. Orientation of the double bond. Reactivity-effects of substrate structures, attacking base, the leaving group and the medium. Mechanism and orientation in pyrolytic elimination.

Unit-V- Pericyclic Reactions

Molecular orbital symmetry, Frontier orbitals of ethylene, 1,3-butadiene, 1,3,5-hexatriene and allyl system. Classification of pericyclic reactions. Woodward-Hoffmann correlation diagrams. FMO and PMO approach. Electrocyclic reactions-conrotatory and disrotatory motions, $4n$, $4n+2$ and allyl systems. Cycloadditions-antarafacial and suprafacial additions, $4n$ and $4n+2$ systems, $2+2$ addition of ketenes, 1,3 dipolar cycloadditions and cheletropic reactions. Sigmatropic rearrangements-suprafacial and antarafacial shifts of H, sigmatropic involving carbon moieties, 3,3- and 5,5 sigmatropic rearrangements. Claisen Cope and Aza-Cope rearrangements. Fluxional tautomerism. Ene reaction.

Scheme of Question Paper

Type of Question	No. of Question	Marks Alloted	Total marks
Objective	10	1	10
Short Answer	5	5	25
Long Answer	5	10	50
Total			85

Book Suggested

1. Advanced Organic Chemistry-Reactions, Mechanism and Structure, Jerry March, John Wiley.
2. Advanced Organic Chemistry, F.A. Carey and R.J. Sundberg, Plenum.
3. A Guide Book to Mechanism in Organic Chemistry, Peter Sykes, Longman.
4. Structure and Mechanism in Organic Chemistry, C.K. Ingold, Cornell University Press.
5. Organic Chemistry, R.T. Morrison and R.N. Boyd, Prentice-Hall.
6. Modern Organic Reactions, H.O. House, Benjamin.
7. Principles of Organic Synthesis, R.O.C. Norman and J.M. Coxon, Blackie Academic &* Professionalsl.
8. Reaction Mechanism in Organic Chemistry, S.M. Mukherji and S.P. Singh, Macmillan.
9. Pericyclic Reactions, S.M. Mukherji, Macmillan, India
10. Stereochemistry of Organic Compounds, D.Nasipuri, New Age International.
11. Stereochemistry of Organic Compounds, P.S. Kalsi, New Age International.

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Department of Higher Education, Govt. of M.P.

Post Graduate Semester Wise Syllabus

As recommended by Central Board of Studies and approved by the Governor of M. P.
(2023-24)

Max Marks	: 100 (Theory 85+ CCE 15)
Class/Semester	: M.Sc. II
Subject	: Chemistry
Paper	: III (Compulsory)
Title	: PHYSICAL CHEMISTRY II

Unit-I- Chemical Dynamics

Methods of determining rate laws, collision theory of reaction rates, steric factor, activated complex theory, arrhenius equation and the activated complex theory; ionic reactions, kinetic salt effects, steady state kinetics, kinetic and thermodynamic control of reactions, treatment of unimolecular reactions. Dynamic chain (hydrogen-bromine reaction, pyrolysis of acetaldehyde, decomposition of ethane), photochemical (hydrogen-bromine and hydrogen-chlorine reactions) and homogenous catalysis, kinetics of enzyme reactions, general features of fast reactions, study of fast reactions by flow method, relaxation method, flash photolysis of the nuclear magnetic resonance method, dynamics of unimolecular reaction (Lindemann Hinshelwood and Rice-Ramsperger-Kassel-Marcus (RRKM) theories for unimolecular reactions).

Unit-II - Surface Chemistry

Adsorption

Surface tension, capillary action, pressure difference across curved surface (Laplace equation), vapour pressure of droplets (Kelvin equation), Gibbs adsorption isotherm, estimation of surface area (BET equation), Surface films on liquids (Electro-kinetic phenomenon).

Micelles

Surface active agents, classification of surface active agents, micellization, hydrophobic interaction, critical micellar concentration (CMC), factors affecting the CMC of surfactants, counter ion binding to micelles, thermodynamics of micellization-phase separation and mass action models, solubilization, micro emulsion, reverse micelles.

Unit-III - Macromolecules

Polymer-definition, types of polymers, electrically conducting, fire resistant, liquid crystal polymers, kinetics of polymerization, mechanism of polymerization. Molecular mass, number and mass average molecular mass, molecular mass determination (Osmometry, viscometry, diffusion and light scattering methods), sedimentation, chain configuration of macromolecules, calculation of average dimension of various chain structures.

Unit-IV- Non Equilibrium Thermodynamics

Thermodynamic criteria for non-equilibrium states, entropy production and entropy flow, entropy balance equations for different irreversible processes (e.g., heat flow, chemical reaction etc.) transformations of the generalized fluxes and forces, non equilibrium stationary states.

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phenomenological equations, microscopic reversibility and Onsager's reciprocity relations, electrokinetic phenomena, diffusion, electric conduction.

Unit-V

Electrochemistry

Electrochemistry of solutions. Debye-Huckel-Onsager treatment and its extension, ion solvent interactions. Debye-Huckel-Jerum mode. Thermodynamics of electrified interface equations. Derivation of electro capillarity, Lippmann equations (surface excess), methods of determination. Structure of electrified interfaces. Overpotentials, exchange current density, derivation of Butler Volmer equation, Tafel plot. Quantum aspects of charge transfer at electrodes-solution interfaces, quantization of charge transfer, tunneling. Semiconductor interfaces-theory of double layer at semiconductor, electrolyte - solution interfaces, structure of double layer interfaces. Effect of light at semiconductor solution interface. Polarography theory, Ilkovic equation; half wave potential and its significance.

Scheme of Question Paper

Type of Question	No. of Question	Marks Alloted	Total marks
Objective	10	1	10
Short Answer	5	5	25
Long Answer	5	10	50
Total			85

Books Suggested

1. Physical Chemistry, P.W. Atkins, ELBS.
2. Introduction to Quantum Chemistry, A.K. Chandra, Tata Mc Graw Hill.
3. Quantum Chemistry, Ira N. Levine, Prentice Hall.
4. Coulson's Valence, R. Mc Ween y, ELBS.
5. Chemical Kinetics. K.J. Laidler, McGraw-Hill.
6. Kineties and Mechanism of Chemical Transformation J.Rajaraman and J.Kuriacose, Mc Millan.
7. Micelles, Theoretical and Applied Aspects, V. MOraoi, Plenum.
8. Modern Electrochemistry Vol. I and Vol II J.O.M. Bockris and A.K.N. Reddy, Planum.
9. Introduction to Polymer Science, V.R. Gowarikar, N.V. Vishwanathan and J.Sridhar, Wiley Eastern.

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Post Graduate Semester Wise Syllabus

As recommended by Central Board of Studies and approved by the Governor of M. P.
(2023-24)

Max Marks : 100 (Theory 85+ CCE 15)
Class/Semester : M.Sc. II
Subject : Chemistry
Paper : IV (Compulsory)
Title : Spectroscopy II and Diffraction Methods

Unit-I - Nuclear Magnetic Resonance Spectroscopy

Nuclear spin, nuclear resonance, saturation, shielding of magnetic nuclei, chemical shift and its measurements, factors influencing chemical shift, deshielding, spin-spin interactions, factors influencing coupling constant "j" Classification (AXB, AMX, ABC, A2B2 etc.). spin decoupling; basic ideas about instrument, NMR studies of nuclei other than proton-¹³C, ¹⁹F and ³¹P. FT NMR, advantages of FT NMR.

Unit II- Nuclear Quadrupole Resonance Spectroscopy

Quadrupole nuclei, quadrupole moments, electric field gradient, coupling constant, splitting. Applications.

Unit-III- Electron Spin Resonance Spectroscopy

Basic principles, zero field splitting and Kramer's degeneracy, factors affecting the 'g' value. Isotropic and anisotropic hyperfine coupling constants, spin Hamiltonian, spin densities and McConnell relationship, measurement techniques, applications.

Unit-IV- X-ray Diffraction

Bragg condition, Miller indices, Laue Method, Bragg method, Debye Scherrer method of X-ray structural analysis of crystals, index reflections, identification of unit cells from systematic absences in diffraction pattern, Structure of simple lattices and X-ray intensities, structure factor and its relation to intensity and electron density, phase problem. Description of the procedure for an X-ray structure analysis, absolute configuration of molecules.

Unit-V- Electron Diffraction

Scattering intensity vs. scattering angle. Wierl equation, measurement technique, elucidation of structure of simple gas phase molecules. Low energy electron diffraction and structure of surfaces.

Neutron Diffraction Scattering of neutrons by solids measurement techniques, Elucidation of structure of magnetically ordered unit cells.

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Scheme of Question Paper

Type of Question	No. of Question	Marks Alloted	Total marks
Objective	10	1	10
Short Answer	5	5	25
Long Answer	5	10	50
Total			85

Books suggested

1. Modern Spectroscopy, J.M. Hollas, John Wiley.
2. Applied Electron Spectroscopy for chemical analysis d. H. Windawi and F.L. Ho, Wiley Interscience.
3. NMR, NQR, EPr and Mossbauer Spectroscopy in Inorganic Chemistry, R.V.Parish, Ellis Harwood.
4. Physical Methods in Chemistry, R.S. Drago, Saunders College.
5. Chemical Applications of Group Theory, F.A. Cotton.
6. Introduction to Molecular Spectroscopy, G.M. Barrow, Mc Graw Hill.
7. Basic Principles of Spectroscopy, R. Chang, Mc Graw Hill.
8. Theory and Application of UV Spectroscopy, H.H. Jaffe and M. Orchin, IBHOxford.
9. Introduction to Photoelectron Spectroscopy, P.K. Ghosh, John Wiley.
10. Introduction to Magnetic Resonance. A Carrington and A.D. Maclachalan, harper & Row.

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Post Graduate Semester Wise Syllabus

As recommended by Central Board of Studies and approved by the Governor of M. P.
(2023-24)

Max Marks : 100 (Theory 85+ CCE 15)
Class/Semester : M.Sc. II
Subject : Chemistry
Paper : V (Compulsory)
Title : COMPUTERS FOR CHEMISTS

This is a theory cum-laboratory course with more emphasis on laboratory work.

Unit-I- Introduction to computers and Computing

Basic structure and functioning of computer with a PC as illustrative example. Memory I/O devices. Secondary storage Computer languages. Operating systems with DOS as an example Introduction to UNIX and WINDOWS. Principles of programming Algorithms and flow-charts.

Unit-II- Computer Programming in FORTRAN/C/BASIC

(the language features are listed here with reference to FORTRAN. The instructor may choose another language such as BASIC or C the features may be replaced appropriately). Elements of the computer language. Constants and variables. Operations and symbols Expressions. Arithmetic assignment statement. Input and output Format statement. Termination statements. Branching statements as IF or GO TO statement. LOGICAL variables. Double precision variables. Subscripted variables and DIMENSION. DO statement FUNCTION AND SUBROUTINE. COMMON and DATA statement (Student learn the programming logic and these language features by hands on experience on a personal computer from the beginning of this topic.)

Unit-III- Programming in Chemistry

Developing of small computer codes using any one of the languages FORTRAN/C/BASIC involving simple formulae in Chemistry, such as Vander Waals equation. Chemical kinetics (determination of Rate constant) Radioactive decay (Half Life and Average Life). Determination Normality, Molarity and Molality of solutions. Evaluation Electronegativity of atom and Lattice Energy from experimental determination of molecular weight and percentage of element organic compounds using data from experimental metal representation of molecules in terms of elementary structural features such as bond lengths, bond angles.

Unit-IV- Use of Computer programmes

Operation of PC. Data Processing. Running of standard Programs and Packages such as MS WORD, MS EXCEL -special emphasis on calculations and chart formations. X-Y plot. Simpson's Numerical Integration method. Programmes with data preferably from physical chemistry laboratory.

Unit V- Internet

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Application of Internet for Chemistry with search engines, various types of files like PDF, JPG, RTF and Bitmap. Scanning, OMR, Web camera.

Scheme of Question Paper

Type of Question	No. of Question	Marks Alloted	Total marks
Objective	10	1	10
Short Answer	5	5	25
Long Answer	5	10	50
Total			85

Book Suggested :

1. Fundamentals of Computer : V. Rajaraman (Prentice Hall)
2. Computers in Chemistry : K.V. Raman (Tata Mc Graw Hill)
3. Computer Programming in FORTRAN IV-V Rajaraman (Prentice Hall)

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Post Graduate Semester Wise Syllabus

As recommended by Central Board of Studies and approved by the Governor of M. P.
(2023-24)

Class & Semester : M.Sc. II
Subject : Chemistry
Paper : Chemistry Practical
Title : Inorganic Chemistry

Max Marks : 33

Time: 6-8hours

Quantitative analysis	8
Chromatography	8
Preparation	8
Record	04
Viva Voce	05
TOTAL	33

Quantitative analysis

Separation and determination of two metal ions- Cu-Ni, Ni-Zn, Cu- Fe, etc involving volumetric and gravimetric methods.

Chromatography

Separation of cations and anions by Column Chromatography : Ion exchange.

Preparations

Preparation of selected inorganic compounds and their studies by I.R. electronic spectra, Mossbauer, E.S.R. and magnetic susceptibility measurements. Handling of air and moisture sensitive compounds.

1. $[\text{Co}(\text{NH}_3)_6][\text{Co}(\text{NO}_2)_6]$
2. cis- $[\text{Co}(\text{trien})(\text{NO}_2)_2]\text{Cl}\cdot\text{H}_2\text{O}$
3. $\text{Hg}[\text{Co}(\text{SCN})_4]$
4. $[\text{Co}(\text{Py})_2\text{Cl}_2]$
5. $[\text{Ni}(\text{NH}_3)_6]\text{Cl}_2$
6. $\text{Ni}(\text{dmg})_2$
7. $[\text{Cu}(\text{NH}_3)_4]\text{SO}_4\cdot\text{H}_2\text{O}$

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Other smaller signatures in blue ink are scattered around the list items, including one near item 1, one near item 2, one near item 3, one near item 4, one near item 5, one near item 6, and one near item 7.

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Post Graduate Semester Wise Syllabus

As recommended by Central Board of Studies and approved by the Governor of M. P.
(2023-24)

Class & Semester : M.Sc. II
Subject : Chemistry
Paper : Chemistry Practical
Title : Organic Chemistry

Max Marks : 33

Time: 6-8hours

Organic Synthesis	12
Quantitative Analysis	12
Record	04
Viva Voce	05
TOTAL:	33

Organic Synthesis

Synthesis of p-nitroaniline and p-bromoaniline. The products may be characterized by spectral techniques.

Synthesis involving following reactions

1. Aromatic electrophilic substitution
2. Condensation reactions
3. Rearrangement reactions
4. Diazotization reactions

Quantitative Analysis

Determination of the percentage or number of hydroxyl groups in an organic compound by acetylation method. Estimation of amines/phenols using bromate bromide solution/or acetylation method. Determination of iodine and Saponification values of an oil sample. Determination of DO, COD and BOD of water sample.

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Post Graduate Semester Wise Syllabus

As recommended by Central Board of Studies and approved by the Governor of M. P.
(2023-24)

Class & Semester : M.Sc. II
Subject : Chemistry
Paper : Chemistry Practical
Title : Physical Chemistry

Max Marks : 34

Time: 6-8hours

Chemical Kinetics	8
Potentiometry / pH metry	9
Polarimetry	8
Record	4
Viva Voce	5

Chemical Kinetics

- Determination of the effect of (a) Change of temperature (b) Change of concentration of reactant and catalyst and (c) Ionic strength of the media on the velocity constant of hydrolysis of an ester/ionic reaction.
- Determination of the velocity constant of hydrolysis of an ester/ionic reaction in micellar media.
- Determination of the velocity constant for the oxidation of iodide ions by hydrogen peroxide study the kinetics as an iodine clock reactions
- Flowing clock reactions (Ref : Experiments in Physical Chemistry by Showmaker)
- Determination of the primary salt effect on the kinetics of ionic reaction and testing of the Bronsted relationship (iodide ion is oxidised by persulphate ion).
- Oscillatory reaction.

Potentiometry/pH metry

- Determination of strengths of halides in a mixture potentiometrically.
- Determination of the valency of mercurous ions potentiometrically.
- Determination of the strength of strong and weak acids in a given mixture using a potentiometer/pH meter.
- Determination of temperature dependence of EMF of a cell.
- Determination of the formation constant of silver-ammonia complex and stoichiometry of the complex potentiometrically.
- Acid-base titration in a non-aqueous media using a pH meter.
- Determination of activity and activity coefficient of electrolytes.
- Determination of the dissociation constant of acetic acid in DMSO, DMF, acetone and dioxane by titrating it with KOH.

9. Determination of the dissociation constant of monobasic/dibasic acid by Albert-Sderjeant method.

10. Determination of thermodynamic constants, DG, DS, and DH for the reaction by e.m.f. method. $Zn + H_2SO_4 \rightarrow ZnSO_4 + 2 H$

Polarimetry

1. Determination of rate constant for hydrolysis/inversion of sugar using a polarimeter.

2. Enzyme kinetics-inversion of sucrose.

Books Suggested

11. Vogel's Textbook of Quantitative Analysis, revised, J. Bassett, R.C. Denney, G.H. Jeffery and J. Mendham, ELBS.

12. Synthesis and Characterization of Inorganic Compounds, W.L. Jolly, Prentice Hall.

13. Experiments and Techniques in Organic Chemistry, D.P. Pasto, C. Johnson and M. Miller, Prentice Hall.

14. Macroscale and Microscale Organic Experiments, K.L. Williamson, D.C. Heath.

15. Systematic Qualitative Organic Analysis, H. Middleton, Adward Arnold.

16. Handbook of Organic Analysis-qualitative and Quantitative. H. Clark, Adward Arnold.

17. Vogel's Textbook of Practical Organic Chemistry, A.R. Tatchell, John Wiley.

18. Practical Physical Chemistry, A.M. James and F.E. Prichard, Longman.

19. Findley's Practical Physical chemistry, B.P. Levitt, Longman.

20. Experimental Physical Chemistry, R.C. Das and B. Behera, Tata McGraw Hill.

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(NAAC Accredited 'A' Grade)



SYLLABUS

PG

SESSION- 2023-24

Class: M.Sc. Semester-III & IV

SUBJECT: Chemistry

Shr *Muro* *Q* *Mita Dutta Koush* *Vaish* *Koush* *Shr*

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(2023-24)

Scheme of Marks Distribution
M. Sc Sem III

Maximum Marks :100
Theory:85
CCE-15

S.No.	Paper Code	Subject	Paper	Max Marks
1.	MCH-	Application of spectroscopy -I	Compulsory	85
2.	MCH-	Photochemistry	Compulsory	85
3.	MCH-	Environment Chemistry	Compulsory	85
4.	MCH-	Polymers	Optional-I	85
5.	MCH-	Industrial Chemistry-Heavy Chemicals and petroleum	Optional-II	85

Scheme of Question Paper

Type of Question	No. of Question	Marks Allotted	Total marks
Objective	10	1	10
Short Answer	5	5	25
Long Answer	5	10	50
Total			85

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 (2023-24)

Max Marks : 100 (Theory 85+ CCE 15)
Class & Semester : M.Sc. III
Subject : Chemistry
Paper : I (Compulsory)
Title : APPLICATION OF SPECTROSCOPY-I

Unit-I

Electronic Spectroscopy: Electronic Spectral Studies for $d^1 - d^9$ systems in octahedral, tetrahedral and square planer complexes.

Unit II

Vibrational Spectroscopy: Symmetry and shapes of AB_2 , AB_3 , AB_4 , AB_5 and AB_6 , mode of bonding of ambidentate ligands, nitrosyl, ethylenediamine and diketonato complexes, application of resonance Raman spectroscopy and applications of vibrational spectroscopy.

Unit III

Nuclear Magnetic Resonance Spectroscopy I: General introduction and definition, chemical shift, spin-spin interaction, shielding and deshielding mechanism, mechanism of measurement of chemical shift values and correlation for protons bonded to carbon (aliphatic, olefinic, aldehydic and aromatic) and other nuclei (alcohols, phenols, enols, carboxylic acids, amines, amides and mercapto).

Unit IV

Nuclear Magnetic Resonance Spectroscopy II: Chemical exchange, effect of deuteration, complex spin-spin interactions between two, three, four and five nuclei (1 order spectra), stereochemistry, hindered rotation, Karplus curve variation of coupling constant with disordered angle, NMR shift reagents, solvent effects, Nuclear overhauser effect (NOE).

Unit V

Mossbauer Spectroscopy: Basic principles, spectral parameters and spectrum display. Application of the technique to the studies of (1) bonding and structures of Fe^{+2} and Fe^{+3} compounds including those of intermediate spin, (2) Sn^{+2} and Sn^{+4} compounds nature of M-L bond, coordination number, structure and (3) detection of oxidation state and inequivalent MB atoms.

Scheme of Question Paper

Type of Question	No. of Question	Marks Alloted	Total marks
Objective	10	1	10
Short Answer	5	5	25
Long Answer	5	10	50
Total			85

Book Suggested:

- Physical Methods for Chemistry, R.S. Drago, Saunders Company.
- Structural Methods in Inorganic Chemistry, E.A.V. Ebsworth, D.W.H. Rankin and S. Cradock, ELBS.
- Infrared and Raman Spectral: Inorganic and Coordination Compounds K. Nakamoto, Wiley.

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4. Progress in Inorganic Chemistry vol., 8, ed., F.A. Cotton, vol., 15 ed. S.J.Lippard, Wiley.
5. Transition Metal Chemistry ed. R.L. Carlin vol. 3 dekker.
6. Inorganic Electronic Spectroscopy, A.P.B. Lever, Elsevier.
7. NMR, NQR, EPR and Mossbauer Spectroscopy in Inorganic Chemistry, .V.Parish, Ellis Haywood.
8. Practical NMR Spectroscopy, M.L. Martin. J.J. Deepish and G.J. Martin, Heyden.
9. Spectrometric Identification of Organic Compounds, R.M. Silverstein, G.C.Bassler and T.C. Morrill, John Wiley.
10. Introduction to NMR spectroscopy, R.J. Abraham, J. Fisher and P. Loftus, Wiley.
11. Application of Spectroscopy of Organic Compounds, J.R. Dyer Prentice Hall.
12. Spectroscopic Methods in Organic Chemistry D.H. Williams, I. Fleming, Tata McGraw-Hill.
13. Structural Methods in Inorganic Chemistry, E.A.V. Ebsworth, D.W.H. Rankin and S. Cradock, ELBS.
14. Introduction to NMR spectroscopy, R.J. Abraham, J. Fisher and P. Loftus, Wiley.

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Post Graduate Semester Wise Syllabus

As recommended by Central Board of Studies and approved by the Governor of M. P.
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Max Marks : 100 (Theory 85+ CCE 15)
Class/Semester : M.Sc. III
Subject : Chemistry
Paper : II (Compulsory)
Title : PHOTOCHEMISTRY

Unit-I- Photochemical Reactions

Interaction of electromagnetic radiation with matter, types of excitations, fate of excited molecule, quantum yield, transfer of excitation energy, actinometry.

Unit –II- Determination of Reaction Mechanism

Classification, rate constants and life times of reactive energy state, determination of rate constants of reactions. Effect of light intensity on the rate of photochemical reactions. Types of photochemical reactions-photo dissociation, gas-phase photolysis.

Unit-III- Photochemistry of Alkenes

Intramolecular reactions of the olefinic bond-geometrical isomerism, cyclisation reactions, rearrangement of 1,4- and 1,5-dienes.

Photochemistry of Aromatic Compounds

Isomerization, additions and substitutions.

Unit-IV- Photochemistry of Carbonyl Compounds

Intramolecular reactions of carbonyl compounds-saturated, cyclic and acyclic, β,γ unsaturated and α,β unsaturated compounds, cyclohexadienones. Intermolecular cycloaddition reactions-dimerisations and oxetane formation.

Unit-V- Miscellaneous Photochemical Reactions.

Photo-Fries reactions of annilides, Photo-Fries rearrangement. Barton reaction. Singlet molecular oxygen and its reactions. Photochemical formation of smog. Photodegradation of polymers. Photochemistry of vision.

Scheme of Question Paper

Type of Question	No. of Question	Marks Alloted	Total marks
Objective	10	1	10
Short Answer	5	5	25
Long Answer	5	10	50
Total			85

Books Suggested

1. Fundamentals of photochemistry, K.K. Rothagi-Mukheriji, Wiley-Eastern.
2. Essentials of Molecular Photochemistry, A Gilbert and J. Baggott, Blackwell Scientific Publication.
3. Molecular Photochemistry, N.J. Turro, W.A. Benjamin.
4. Introductory Photochemistry, A. Cox and t. Camp, McGraw Hill.
5. Photochemistry, R.P. Kundall and A. Gilbert. Thomson Nelson.
6. Organic Photochemistry, J. Coxon and B.halton, Cambridge University Press.

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Department of Higher Education, Govt. of M.P.
Post Graduate Semester Wise Syllabus
As recommended by Central Board of Studies and approved by the Governor of M. P.
(2023-24)

Max Marks	: 100 (Theory 85+ CCE 15)
Class & Semester	: M.Sc. III
Subject	: Chemistry
Paper	: III (Compulsory)
Title	: ENVIRONMENTAL CHEMISTRY

Unit-I- Atmosphere

Atmospheric layers, Vertical temperature profile, heat/radiation budget of the earth atmosphere systems. Properties of troposphere, thermodynamic derivation of lapse rate. Temperature inversion. Calculation of Global mean temperature of the atmosphere. Pressure variation in atmosphere and scale height. Biogeochemical cycles of carbon, nitrogen, sulphur, phosphorus, oxygen. Residence times.

Atmospheric Chemistry

Sources of trace atmospheric constituents : nitrogen oxides, sulphur dioxide and other sulphur compounds, carbon oxides, chlorofluorocarbons and other halogen compounds, methane and other hydrocarbons.

Tropospheric Photochemistry

Mechanism of Photochemical decomposition of NO_2 and formation of ozone. Formation of oxygen atoms, hydroxyl, hydroperoxy and organic radicals and hydrogen peroxide. Reactions of hydroxyl radicals with methane and other organic compounds. Reaction of OH radicals with SO_2 and NO_2 . Formation of Nitrate radical and its reactions. Photochemical smog meteorological conditions and chemistry of its formation.

Unit-II - Air Pollution

Air pollutants and their classifications. Aerosols-sources, size distribution and effect on visibility, climate and health.

Acid Rain

Definition, Acid rain precursors and their aqueous and gas phase atmospheric oxidation reactions. Damaging effects on aquatic life, plants, buildings and health. Monitoring of SO_2 and NO_2 . Acid rain control strategies.

Stratospheric Ozone Depletion

Mechanism of Ozone formation, Mechanism of catalytic Ozone depletion, Discovery of Antarctic Ozone hole and Role of chemistry and meteorology. Control Strategies.

Green House Effect

Terrestrial and solar radiation Spectra, Major green house gases and their sources and Global warming potentials. Climate change and consequences.

Urban Air Pollution- Exhaust emissions, damaging effects of carbon monoxide. Monitoring of CO. Control strategies.

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Unit-III- Aquatic Chemistry and Water Pollution

Redox chemistry in natural waters. Dissolved oxygen, biological oxygen demand, chemical oxygen demand, determination of DO, BOD and COD. Aerobic and anaerobic reactions of organic sulphur and nitrogen compounds in water, acid-base chemistry of fresh water and sea water. Aluminum, nitrate and fluoride in water. Petrification. Sources of water pollution. Treatment of waste and sewage. Purification of drinking water, techniques of purification and disinfection.

Unit IV - Environmental Toxicology

Toxic heavy metals : Mercury, lead, arsenic and cadmium. Causes of toxicity.

Bioaccumulation, sources of heavy metals. Chemical speciation of Hg, Pb, As, and Cd. Biochemical and damaging effects.

Toxic Organic Compound : Pesticides, classification, properties and uses of organochlorine and ionospheres pesticides detection and damaging effects.

Polychlorinated biphenyls : Properties, use and environmental continuation and effects.

Polynuclear Aromatic Hydrocarbons : Source, structures and as pollutants.

Unit-V - Soil and Environmental Disasters

Soil composition, micro and macronutrients, soil pollution by fertilizers, plastic and metals. Methods of re-remediation of soil. Bhopal gas tragedy, Chernobyl, three mile island, Minamata Disease, Saveso (Italy), London smog.

Scheme of Question Paper

Type of Question	No. of Question	Marks Alloted	Total marks
Objective	10	1	10
Short Answer	5	5	25
Long Answer	5	10	50
Total			85

Books Suggested

1. Environmental Chemistry, Colin Baird, W.H. Freeman Co. New York, 1998.
2. Chemistry of Atmospheres, R.P. Wayne, Oxford.
3. Environment Chemistry, A.K. De, Wiley Eastern, 2004.
4. Environmental Chemistry, S.E. Manahan, Lewis Publishers.
5. Introduction to atmospheric Chemistry, P.V. Hobbs, Cambridge.

Handwritten signatures and names in blue ink:
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A signature in the center.
A signature on the right with the name "Anita Sultkhan" written below it.
A signature on the far right with the name "Vasika" written above it.
A signature at the bottom right with the name "Anu" written below it.

Sri Sathya Sai College for Women, Bhopal
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Post Graduate Semester Wise Syllabus

As recommended by Central Board of Studies and approved by the Governor of M. P.
 (2023-24)

Max Marks : 100 (Theory 85+ CCE 15)
Class & Semester : M.Sc. III
Subject : Chemistry
Paper : IV (Optional I)
Title : Polymers

Unit-I : Basics

Importance of polymers. Basic concepts : Monomers, repeat units, degree of polymerization Linear, branched and network polymers. Classification of polymers. Polymerization : condensation, addition/radical chain-ionic and co-ordination and copolymerization. Polymerization conditions and polymer reactions. Polymerization in homogeneous and heterogeneous systems.

Unit-II : Polymer Characterization: Polydispersion-average molecular weight concept. Number, weight and viscosity, average molecular weights. Polydispersity and molecular weight distribution. The practical significance of molecular weight. Measurement of molecular-weights. End-group, viscosity, light scattering, osmotic and ultracentrifugation methods, spectroscopic methods, X-ray diffraction study. Microscopy, Thermal analysis.

Unit-III: Analysis and testing of polymers

Chemical analysis of polymers, mechanical testing-tensile strength, flexural strength. Fatigue impact, tear resistance. Hardness and abrasion resistance. Thermal testing including mp, HDT, VSP, Tg low temperature brittleness etc.

Unit-IV: Inorganic Polymers

A general survey and scope of Inorganic Polymers special characteristics, classification, homo and hetero atomic polymers. Structure, Properties and Applications of

- a. Polymers based on boron-borazines, boranes and carboranes.
- b. Polymers based on Silicon, silicone's polymetalloxanes and polymetallosiloxanes, silazanes.

Unit V: Structure, Properties and Application of Polymers

- a. Polymers based on Phosphorous-Phosphazenes, Polyphosphates
- b. Polymers based on Sulphur - Tetrasulphur tetranitride and related compounds.
- c. Co-ordination and metal chelate polymers.

Scheme of Question Paper

Type of Question	No. of Question	Marks Alloted	Total marks
Objective	10	1	10
Short Answer	5	5	25
Long Answer	5	10	50
Total			85

Books Suggested

1. Inorganic Chemistry, J.E. Huheey, Harper Row.
2. Developments in Inorganic polymer Chemistry, M.F. Lappert and G.J. Leigh.
3. Inorganic polymers- N.H. Ray.
4. Inorganic polymers, Graham and Stone.

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5. Inorganic Rings and Cages : D.A. Armitage.
6. Textbook of Polymers Science, F.W. Billmeyer Jr. Wiley.
7. Contemporary Polymer Chemistry, H.R. Alcock and F.W. Lambe, Prentice Hall.
8. Polymer Science, Gowariker, Vishwanathan, Sreedhar.
9. Testing of polymers Vishu Shah.

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Max Marks : 100 (Theory 85+ CCE 15)
Class & Semester : M.Sc. III
Subject : Chemistry
Paper : V(Optional II)
Title : Industrial Chemistry- Heavy Chemicals & Petroleum

Unit I - Water, Gases and Heavy Chemicals

Water: Water Pollutants, their classes with example, Biochemical oxygen demand, thermal pollution, pollution by fertilizer, detergents, pesticides and industrial wastes.

Water Purification : Classical and modern methods – Ion exchange, electro dialysis, Reverse osmosis. Softening of hard water. Chlorination and fluoridation.

Unit II - Gases: Chemistry Large - scale production, storage, hazards and uses of the following industrial gases: Hydrogen, oxygen, nitrogen, carbon dioxide, chlorine, fluorine, sulphur dioxide, phosgene, acetylene, argon, neon and helium .

Heavy Chemicals: Manufacture, physical properties, Analysis, hazard and applications of the following chemicals:

HCl, H₂SO₄, HNO₃, H₃PO₄, polyphosphoric acid .

NaHCO₃, Na₂CO₃, NaOH , NaCl, Na₂S₂O₃, Bleaching Powder, Bromine.

Unit III- Coal and Petroleum

Coal: Origin and economic importance of coal. Coal composition. Coal carbonization. Coal gasification. Coal Gas. Water Gas. Producer gas. Coal tar. Industry and manufacture of coal tar based chemicals and their importance. Role as carcinogens, non-fuel uses of coal and Cl chemistry based on MeOH, CO, CO₂, CH₄ and CH₂O.

Unit IV- Petroleum : Origin and composition, Refining, Reforming Fractionation, Cracking, Knocking and Octane number, Kerosene and Napthe; Liquified Petroleum Gas (L.P.G.), Synthetic Gas, Synthetic Petrol. Petrochemicals, Manufacture of ethylene, Propylene, Butadiene, Xylenes etc. Economic importance with particular reference to India.

Unit V- Fats and Oils - Fats and Oils Natural fats, Edible and Industrial Oils of vegetable origin common fatty acids and glycerides, hydrogenation of unsaturated oils, manufacture of Vanaspati and margarine.

Scheme of Question Paper:

Type of Question	No. of Question	Marks Alloted	Total marks
Objective	10	1	10
Short Answer	5	5	25
Long Answer	5	10	50
Total			85

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Post Graduate Semester Wise Syllabus

As recommended by Central Board of Studies and approved by the Governor of M. P.
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Class & Semester : M.Sc. III
Subject : Chemistry
Paper : Chemistry Practical
Title : Inorganic Chemistry

Max Marks : 33

Time: 6-8 hours

Quantitative determinations of a three component mixture	12
Chromatographic Separations	12
Record	04
Viva Voce	05
Total	33

1. Quantitative determinations of a three component mixture :

One Volumetrically and two gravimetrically

- a. Cu^{+2} , Ni^{+2} , Zn^{+2}
- b. Cu^{+2} , Ni^{+2} , Ag^{+1}

2. Chromatographic Separations

- a. Cadmium and zinc
- b. Zinc and magnesium.
- c. Thin-layer / paper chromatography-separation of nickel, manganese, cobalt and zinc.
Determination of R_f values.

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As recommended by Central Board of Studies and approved by the Governor of M. P.
(2023-24)

Class & Semester : M.Sc. III
Subject : Chemistry
Paper : Chemistry Practical
Title : Organic Chemistry

Max Marks : 33

Time: 6-8hours

Estimations	12
Paper Chromatography	12
Record	04
Viva Voce	05
Total	33

1. Estimation

An Estimation of element: Estimation of Sulphur, Nitrogen and Halogens.

B Estimation of Functional groups and simple organic compounds, Estimation of Carboxylic group, Aldehyde group, Ketonic group, Carbohydrate group etc.

2. Paper Chromatography

Separation and identification of the sugars present in the given mixture of glucose, fructose and sucrose by paper chromatography and determination of R_f values.

3. Polymer Synthesis

Synthesis of some common polymers and their characterization.

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As recommended by Central Board of Studies and approved by the Governor of M. P.
(2023-24)

Class & Semester : M.Sc. III
Subject : Chemistry
Paper : Chemistry Practical
Title : Physical Chemistry

Max Marks : 34

Time: 6-8hours

Spectroscopy	13
Chemical Kinetics	12
Record	04
Viva Voce	05
Total	34

Spectroscopy

- i. Determination of pKa of an indicator (e.g. methyl red) in (a) aqueous and (b) micellar media.
- ii. Determination of stoichiometry and stability constant of Ferric isothiocyanate complex ion in solution.
- iii. Determination of rate constant of alkaline bleaching of Malachite green and effect of ionic strength on the rate of reaction.

Chemical Kinetics

- i. Determination of energy and enthalpy of activation in the reaction of KMnO_4 and benzyl alcohol in acid medium.
- ii. Determination of energy of activation and entropy of activation from a single kinetic run.
- iii. Kinetics of an enzyme micellar catalyzed reaction.
- iv. Kinetics of reaction between acetone and iodine spectrophotometrically .
- v. Spectrometric study of oxidation of alcohol by $\text{Ce}(\text{iv})$.
- vi. Oxidation of alcohols by chromic acid – a kinetic study .
- vii. Determination of CMC spectrophotometrically .

Books Suggested

1. Inorganic Experiments, J. Derek Woolings, VCH.
2. Microscale Inorganic Chemistry, Z. Szafran, R.M. Pike and M.M. Singh, Wiley.
3. Practical Inorganic Chemistry, G. Marr and B. W. Rockett, Van Nostrand.
4. The systematic Identification of Organic Compounds, R.L. Shriner and D.Y. Curtin.

The students are expected to visit industries to gain practical knowledge, if possible.

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Scheme of Marks Distribution
M.Sc Sem IV

Maximum Marks:100
Theory:85
CCE-15

S.No.	Paper Code	Subject	Paper	Max Marks
1.	MCH-	Application of spectroscopy -II	Compulsory	85
2.	MCH-	Solid state chemistry	Compulsory	85
3.	MCH-	Biochemistry Chemistry	Compulsory	85
4.	MCH-	Analytical Chemistry	Optional-I	85
5.	MCH-	Medicinal Chemistry	Optional-II	85

Scheme of Question Paper

Type of Question	No. of Question	Marks Alloted	Total marks
Objective	10	1	10
Short Answer	5	5	25
Long Answer	5	10	50
Total			85

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Post Graduate Semester Wise Syllabus

As recommended by Central Board of Studies and approved by the Governor of M. P.
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Max Marks : 100 (Theory 85+ CCE 15)
Class/Semester : M.Sc. IV
Subject : Chemistry
Paper : I (Compulsory)
Title : APPLICATION OF SPECTROSCOPY-II

Unit-I : Ultraviolet and Visible spectroscopy

Various electronic transitions (185-800 nm) Beer-Lambert law, Effect of solvent on electronic transitions, ultraviolet bands for carbonyl compounds, unsaturated carbonyl compounds, dienes, conjugated polyenes, Fieser Woodward rules for conjugated dienes and carbonyl compounds, ultraviolet spectra of aromatic compounds. Steric effect in biphenyls.

Unit II: Infrared Spectroscopy

Characteristic vibrational frequencies of alkanes, alkenes, alkynes, aromatic compounds, alcohols, ethers, phenols and amines. Detailed study of vibrational frequencies of carbonyl compounds (ketones, aldehydes, esters, amides, acids, anhydrides, lactones, lactams and conjugated carbonyl compounds). Effect of hydrogen bonding and solvent effect on vibrational frequencies, overtones, combination bands and fermi resonance.

Unit-III :Nuclear Magnetic Resonance of paramagnetic Substances in Solution :

The contact and pseudo contact shifts, factor affecting nuclear relaxation , some applications including biochemical systems, an overview of NMR of metal nuclide with emphasis on ¹⁹⁵Pt and ¹¹⁹Sn NMR .

Unit-IV:Carbon-13 NMR Spectroscopy

General considerations, chemical shift (aliphatic olefinic , alkyne, aromatic, heteroaromatic and carbonyl carbon), coupling constants. Two dimension NMR spectroscopy-COSY, NOESY, DEPT, HMBC and HMQC techniques.

Unit V :Mass Spectrometry

Introduction ion production EI, CIFD, ESI and FAB, factors affecting fragmentation, ion analysis, ion abundance, Mass spectral fragmentation of organic compounds, common functional groups, molecular ion peak, metastable peak. Mc Lafferty rearrangement. Nitrogen rule. High resolution mass spectrometry. Structure elucidation of simple molecules using UV- visible, IR, NMR and mass spectral techniques.

Scheme of Questions\ Paper :

Type of Question	No. of Question	Marks Alloted	Total marks
Objective	10	1	10
Short Answer	5	5	25
Long Answer	5	10	50
Total			85

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Book Suggested

1. Physical Methods for Chemistry, R.S. Drago, Saunders Company.
2. Structural Methods in Inorganic Chemistry, E.A.V. Ebsworth, D.W.H. Rankin and S. Cradock, ELBS.
3. Infrared and Raman Spectral : Inorganic and Coordination Compounds K. Nakamoto, Wiley.
4. Progress in Inorganic Chemistry vol., 8, ed., F.A. Cotton, vol., 15 ed. S.J. Lippard, Wiley.
5. Transition Metal Chemistry ed. R.L. Carlin vol. 3 dekker.
6. Inorganic Electronic Spectroscopy, A.P.B. Lever, Elsevier.
7. NMR, NQR, EPR and Mossbauer Spectroscopy in Inorganic Chemistry, V. Parish, Ellis Haywood.
8. Practical NMR Spectroscopy, M.L. Martin. J.J. Deepish and G.J. Martin, Heyden.
9. Spectrometric Identification of Organic Compounds, R.M. Silverstein, G.C. Bassler and T.C. Morrill, John Wiley.
10. Introduction to NMR spectroscopy, R.J. Abraham, J. Fisher and P. Loftus, Wiley.
11. Application of Spectroscopy of Organic Compounds, J.R. Dyer Prentice Hall.
12. Spectroscopic Methods in Organic Chemistry D.H. Williams, I. Fleming, Tata McGraw-Hill.
13. Structural Method in Inorganic Chemistry , E.A.V. Ebsworth , D.W.H. Rankin and S. Cradock , ELBS.
14. Introduction of NMR spectroscopy , R.J. Abraham, J. Fisher and P. Loftus, Wiley.

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Post Graduate Semester Wise Syllabus

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Max Marks : 100 (Theory 85+ CCE 15)
Class/Semester : M.Sc. IV
Subject : Chemistry
Paper : II(Compulsory)
Title : SOLID STATE CHEMISTRY

Unit I: Solid State Reactions

General principles, experimental procedure, co-precipitation as a precursory to solid state reactions, kinetics of solid state reactions.

Unit II: Crystal Defects and Non-Stoichiometry

Perfect and imperfect crystals, intrinsic and extrinsic defects-point defects, line and plane defects, vacancies-Schottky defects and Frenkel defects. Thermodynamics of Schottky and Frenkel defect formation, colour centres, non-stoichiometry and defects.

Unit III: Electronic Properties and Band Theory

Metals insulators and semiconductors, electronic structure of solids band theory, band structure of metals, insulators and semiconductors, Intrinsic and extrinsic semiconductors, doping semiconductors, p-n junctions, super conductors. Optical properties-Application of optical and electron microscopy. Magnetic Properties- Classification of materials : Effect of temperature calculation of magnetic moment, mechanism of ferro and antiferromagnetic ordering super exchange.

Unit IV: Organic Solids

Electrically conducting solids. organic charge transfer complex, organic metals, new superconductors.

Unit V: Liquid Crystals:

Types of liquid crystals: Nematic, Smectic, Ferroelectric, Antiferroelectric, Various theories of LC, Liquid crystal display, New materials.

Scheme of Question Paper

Type of Question	No. of Question	Marks Alloted	Total marks
Objective	10	1	10
Short Answer	5	5	25
Long Answer	5	10	50
Total			85

Books Suggested.

1. Solid state chemistry and its applications, A.R. West. Peenum.
2. Principles of the Solid State, H.V. Keer, Wiley Eastern.
3. Solid State Chemistry, N.B. Hannay.
4. Solid State Chemistry, D.K. Chakrabarty, New Wiley Eastern.

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Max Marks : 100 (Theory 85+ CCE 15)
Class/Semester : M.Sc. IV
Subject : Chemistry
Paper : III(Compulsory)
Title : BIOCHEMISTRY

Unit I: Metal Ions in Biological Systems

Bulk and trace metals with special reference to Na, K, Mg, Ca, Fe, Cu, Zn, Co, and K^+/Na^+ pump.

Bioenergetics and ATP Cycle.

DNA polymerisation, glucose storage, metal complexes in transmission of energy; chlorophyll's, photosystem I and photosystem II in cleavage of water.

Transport and Storage of Dioxygen

Haem proteins and oxygen uptake, structure and function of haemoglobin, myoglobin, haemocyanin and hemerythrin, model synthetic complexes of iron, cobalt and copper.

Unit II: Electron Transfer in Biology

Structure and function of metal of proteins in electron transport processes cytochrome's and iron-sulphur proteins, synthetic models.

Nitrogen fixation

Biological nitrogen fixation and its mechanism, nitrogenase, Chemical nitrogen fixation.

Unit III: Enzymes

Introduction and historical perspective, chemical and biological catalysis, remarkable properties of enzymes like catalytic power, specificity and regulation. Nomenclature and classification, extraction and purification. Fischer's lock and key and Koshland's induced fit hypothesis, concept and identification of active site by the use of inhibitors, affinity labeling and enzyme modification by site-directed mutagenesis. Enzyme kinetics, Michael's-Menten and Lineweaver Burk plots, reversible and irreversible inhibition.

Mechanism of Enzyme Action

Transition-state theory, orientation and Steric effect, acid-base catalysis, covalent catalysis, strain or distortion. Examples of some typical enzyme mechanisms for chemotrypsin, ribonuclease, lysozyme and carboxypeptidase.

Kinds of Reactions Catalysed by Enzymes

Nucleophilic displacement on a phosphorus atom, multiple displacement reactions and the coupling of ATP cleavage to endergonic processes. Transfer of sulphate, addition and elimination reactions, enolic intermediates in Isomerisation reactions, β -Cleavage and condensation, some isomerization and rearrangement reactions. Enzyme catalyzed carboxylation and decarboxylation.

Unit IV : Co-Enzyme Chemistry

Cofactors as derived from vitamins, coenzymes, prosthetic groups, apoenzymes. Structure and biological functions of coenzyme A, thiamine pyrophosphate, pyridoxal phosphate, NAD⁺, NADP⁺, FMN, FAD, lipoic acid, vitamin B12. Mechanisms of reactions catalyzed by the above cofactors.

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Enzyme Models

Host-guest chemistry, chiral recognition and catalysis, molecular recognition, molecular asymmetry and prochirality Biometric chemistry, crown ether, cryptates. Cyclodextrins, cyclodextrin-based enzyme models, clixarenes, ionospheres, micelles synthetic enzymes or synzymes.

Biotechnological Applications of Enzymes: Large-scale production and purification of enzymes, techniques and methods of immobilization of enzymes, effect of immobilization on enzyme activity, application of immobilized enzymes, use of enzymes in food and drink industry-brewing and cheese making, syrups from corn starch, enzymes as targets for drug design. Clinical uses of enzymes, enzyme therapy, enzymes and recombinant DNA Technology.

Unit V : Biological Cell and its Constituents

Biological cell, structure and functions of proteins, enzymes, DNA and RNA in living systems. Helix coils transition.

Bioenergetics

Standard free energy change in biochemical reactions, exergonic, endergonic. Hydrolysis of ATP, synthesis of ATP from ADP.

Biopolymer Interactions

Forces involved in biopolymer interactions. Electrostatic charges and molecular expansion, hydrophobic forces, dispersion force interactions. Multiple equilibrium and various types of binding processes in biological systems. Hydrogen ion titration curves.

Cell Membrane and Transport of Ions

Structure and functions of cell membrane, ion transport through cell membrane, irreversible thermodynamic treatment of membrane transport. Nerve conduction.

Scheme of Question Paper

Type of Question	No. of Question	Marks Alloted	Total marks
Objective	10	1	10
Short Answer	5	5	25
Long Answer	5	10	50
Total			85

Book Suggested

1. Principles of Bioinorganic Chemistry, S.J. Lippard and J.M. Berg, University Science Books.
2. Bioinorganic Chemistry, I. Bertini, H.B. Gray, S.J. Lippard and J.S. Valentine, University Science Books.
3. Inorganic biochemistry vol. I and II ed. G.L. Eichhorn, Elsevier.
4. Progress in Inorganic Chemistry, Vol 18 and 38 ed J.J. Lippard, Wiley.
5. Bioorganic Chemistry : A chemical Approach to Enzyme Action, Hermann Dugas and C. Penny, Springer Verlag.
6. Understanding Enzymes, Trevor Palmer, Prentice Hall.
7. Enzyme Chemistry : Impact and applications, Ed. Collin J suckling, chemistry.
8. Enzyme Mechanisms Ed. M.I. Page and A Williams, Royal Society of Chemistry.
9. Fundamentals of Enzymology, N.C. Price and L. Stevens. Oxford University Press.
10. Immobilized Enzymes : An Introduction and Applications in Biotechnology, Michael ID. Trevan, Hohn Wiley.
11. Enzymatic Reaction Mechanisms. C. Walsh. W.H. Freeman.
12. Enzyme Structure and Mechanism, A Fersht, W.H. Freeman
13. Biochemistry : The Chemical Reactions of Living Cells, D.E. Metzler, Academic Press.

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Post Graduate Semester Wise Syllabus

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Max Marks : 100 (Theory 85+ CCE 15)
Class/Semester : M.Sc. IV
Subject : Chemistry
Paper : IV (Optional I)
Title : ANALYTICAL CHEMISTRY

Unit I: Introduction

Role of analytical chemistry Classification of analytical methods classical and instrumental. Types of instrumental analysis. Selecting an analytical method. Neatness and cleanliness. laboratory operations and practices. Analytical balance. Techniques of weighing, errors. Volumetric glassware cleaning and calibration of glassware. Sample Volumetric glassware cleaning and Calibration of glassware. Sample preparation dissolution and decompositions. Gravimetric techniques. Selecting and handling of reagents. Laboratory notebooks. Safety in the analytical laboratory.

Errors and Evaluation

Definition of terms in mean and median. Precision-standard deviation, relative standard deviation. Accuracy-absolute error, relative error. Types of error in experimental data determinate (systematic), indeterminate (or random) and gross. Sources of error and the effects upon the analytical results. Methods for reporting analytical data. Statistical evaluation of data-indeterminate errors. The uses of statistics.

Unit II :Food analysis

Moisture, ash, crude protein, fat crude fiber, carbohydrates, calcium, potassium, sodium and phosphate. Food adulteration-common adulterants in food, contamination of food stuffs. Microscopic examination of foods for adulterants. Pesticide analysis in food products. Extraction and purification of sample. HPLC. Gas chromatography for organophosphates. Thin-layer chromatography for identification of chlorinated pesticides in food products.

Unit-III: Analysis of Water Pollution

Origin of Waste water, types, water pollutants and their effects. Sources of water pollution-domestic, industrial, agricultural soil and radioactive wastes as sources of pollution. Objectives of analysis-parameter for analysis-colour, turbidity, total solids, conductivity, acidity, alkalinity, hardness, chloride, sulphate, fluoride, silica, phosphates and different forms of nitrogen, Heavy metal pollution-public health significance of cadmium, chromium, copper, lead, zinc, manganese, mercury and arsenic. General survey of instrumental technique for the analysis of heavy metals in aqueous systems. Measurements of DO, BOD, and COD. Pesticides as water pollutants and analysis. Water pollution laws and standards.

Unit-IV: Analysis of soil, Fuel, Body Fluids and Drugs

(a) **Analysis of Soil:** moisture, pH, total nitrogen, phosphorus, silica, lime, magnesia, manganese, sulphur and alkali salts.

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Amila Rutt Kovan
Nayana
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Fuel analysis : liquid and gas. Ultimate and proximate analysis-heating values-grading of coal. Liquid fuels-flash point, aniline point, octane number and carbon residue. Gaseous fuels-produced gas and water gas-calorific value.

Unit V: (a)Clinical Chemistry: Composition of blood-collection and preservation of samples. Clinical analysis. Serum electrolytes, blood glucose, blood urea nitrogen, uric acid, albumin, globulins, barbiturates, acid and alkaline phosphates. Immunoassay: principles of radio immunoassay (RIA) and applications. The blood gas analysis trace elements in the body.

(b) Drug analysis: Narcotics and dangerous drug. Classification of drugs. Screening by gas and thin-layer chromatography and spectrophotometric measurements.

Scheme of Question Paper

Type of Question	No. of Question	Marks Allotted	Total marks
Objective	10	1	10
Short Answer	5	5	25
Long Answer	5	10	50
Total			85

Book Suggested

1. Analytical Chemistry, G.D. Christian, J.Wicy.
2. Fundamentals of analytical Chemistry. D.A. Skoog. D.M. West and F.J. Hooler, W.B. Saunders.
3. Analytical Chemistry-Principles. J.H. Kennedy. W.B. Saunders.
4. Analytical Chemistry-Principles and Techniques. LG. Hargis. Prentice Hall.
5. Principles of Instrumental analysis D.A. Skoog and J.L. Loary, W.B. Saunders.
6. Principles of Instrumental Analysis D.A. Skoog W.B. Saunders.
7. Quantitative Analysis, R.A. Day, Jr. and A.L. Underwood, Prentice Hall.
8. Environmental Solution, S.M. Khopkar, Wiley Eastern.
9. Basic Concepts of Analysis Chemistry, S.M. Khopkar, Wiley Eastern.
10. Handbook of Instrumental Techniques for Analytical Chemistry, F. Settle, Prentice Hall.

Handwritten signatures and notes in blue ink:
- A signature that looks like "Kuro".
- A signature that looks like "S.S.R".
- A signature that looks like "S.B".
- A signature that looks like "Vijay".
- A signature that looks like "Anita Subhakar".
- A signature that looks like "Nisha".
- A signature that looks like "Anu".

Sri Sathya Sai College for Women, Bhopal
(An Autonomous College Affiliated to Barkatullah University Bhopal)
Department of Higher Education, Govt. of M.P.
Post Graduate Semester Wise Syllabus

As recommended by Central Board of Studies and approved by the Governor of M. P.
(2023-24)

Max Marks : 100 (Theory 85+ CCE 15)
Class/Semester : M.Sc. IV
Subject : Chemistry
Paper : V(Optional II)
Title : MEDICINAL CHEMISTRY

Unit I : Structure and activity : Relationship between chemical structure and biological activity (SAR). Receptor Site Theory. Approaches to drug design. Introduction to combination synthesis in drug discovery. Factors affecting bioactivity. QSAR-Free-Wilson analysis, Hansch analysis, relationship between Free-Wilson analysis and Hansch analysis.

Unit II : Pharmacodynamics: Introduction, elementary treatment of enzymes stimulation, enzyme inhibition, sulfonamides, membrane active drugs, drug metabolism, xenobiotics, biotransformation, significance of drug metabolism in medicinal chemistry.

Unit III : Antibiotics and antibacterials: Introduction, Antibiotic β -Lactam type - Penicillins, Cephalosporins, Antitubercular, Streptomycin, Broad spectrum antibiotics. Tetracyclines, Anticancer – Dactinomycin (Actinomycin D)

Unit IV :Antifungal: Polyenes, Antibacterial - Ciprofloxacin, Norfloxacin, Antiviral - Acyclovir.
Antimalarials: Chemotherapy of malaria. SAR. Chloroquine, Chloroguanide and Mefloquine

Unit V: Non-steroidal Anti-inflammatory Drugs:
Diclofenac Sodium, Ibuprofen and Netopam
Antihistaminic and antiasthmatic agents:
Terfenadine, Cinnarizine, Salbutamol and Beclomethasone dipropionate.

Scheme of Question Paper

Type of Question	No. of Question	Marks Alloted	Total marks
Objective	10	1	10
Short Answer	5	5	25
Long Answer	5	10	50
Total			85

Handwritten signatures and notes:
Aruna
Dhr
Amila Sutt Kovan
Nanda
Arun

Book Suggested

1. Introduction to medicinal chemistry , A. Gringuage, Wiley-VCH.
2. Wilson and Gisvold's Text book of Organic Medicinal and Pharmaceutical chemistry, Ed Robert F Dorge.
3. An introduction to Drug Design, S.S Pandeya and J.R. Dimmock, New Age International.
4. Burger's Medicinal Chemistry and Drug Discovery, Vol-I (Chapter-9 and Chapter-14), Ed. M.E. Wolff, John Wiley.
5. Goodman and Gilman's Pharmacological bases of Therapeutics, Mc Graw-Hill.
6. The Organic Chemistry of Drug Design and Drug Action , R.B. Silverman , Academic Press.
7. Strategies for Organic Drug synthesis and Design , D. Lednicer , John Wiley.
8. Principles of Medicinal Chemistry W.O. Foye .
9. Medicinal Chemistry : The Role of organic chemist in Drug Research, S.M. Roberts and B.J. Pricer.

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(2023-24)

Class & Semester : M.Sc. IV
Subject : Chemistry
Paper : Chemistry Practical
Title : Inorganic Chemistry

Max Marks : 33

Time: 6-8hours

Preparation	12
Spectrophotometric Determinations/ Flame Photometric Determinations	12
Record	04
Viva Voce	05
Total	33

Preparation

Preparation of selected inorganic compounds and their study by IR, electronic spectra, and magnetic susceptibility measurements. Handling of air and moisture sensitive compounds involving vacuum lines. Selection can be made from the following :

1. Sodium amide. Inorg. Synth., 1946, 2, 128.
2. Atomic absorption analysis of Mg and Ca.
3. Synthesis of trichlorodiphenylantimony (V) hydrate. Inorg. Synths., 1985, 23, 194
4. Sodium tetrathionate $\text{Na}_2\text{S}_4\text{O}_6$.
5. Metal complex of dimethyl sulfoxide : $\text{CuCl}_2 \cdot 2\text{DMSO}$ J.Chem. Educ., 1982, 59, 57.
6. Synthesis of metal acetylacetonate : Inorg. Synths, 1957, 5, 130, 1963, 1, 183.
7. Cis and Trans $[\text{Co}(\text{en})_2\text{Cl}_2]^+$.
8. Determination of Cr (III) complex. $[\text{Cr}(\text{H}_2\text{O})_6]\text{NO}_3 \cdot 3\text{H}_2\text{O}$,. Inorg. synths., 1972, 13, 184.
9. Preparation and use of Ferrocene. J. Chem. Edu. 1966, 43, 73; 1976, 53, 730.
10. Preparation of $[\text{Co}(\text{phenanthroline-5,6 quinone})]$.

Spectrophotometric Determinations

- a. Manganese/Chromium in steel sample.
- b. Nickel by extractive spectrophotometric method.
- c. Fluoride/nitrite/phosphate.
- d. Copper-Ethylene diamine complex: Slope-ratio method.

Flame Photometric Determinations

- a. Sodium and potassium when present together.
- b. Lithium/calcium/barium/strontium.
- c. Cadmium and magnesium in tap water

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Dr. J. S. ...
Amrita Sura Kaur
Vaibha
Dr.

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(2023-24)

Class & Semester : M.Sc. IV
Subject : Chemistry
Paper : Chemistry Practical
Title : Organic Chemistry

Max Marks : 33

Time: 6-8hours

Multi- step Synthesis of Organic Compound	12
Spectroscopy/ Spectrophotometric Determinations	12
Record	04
Viva Voce	05
Total	33

Multi- step synthesis of organic compound

The exercise should illustrate the use of organic reagent and may involve purification of the product by chromatographic techniques. Photochemical reaction Benzophenone \rightarrow Benzopinacol \rightarrow Benzopinacolone.

Backmann rearrangement: Benzanilide from benzene: Benzene \rightarrow Benzophenone \rightarrow Benzophenone oxime \rightarrow Benzanilide

Benzilic acid rearrangement: Benzilic acid from benzoin (Benzoin \rightarrow Benzil \rightarrow Benzilic acid) - Synthesis of heterocyclic compounds.

Skraup synthesis: Preparation of quinoline from aniline

Fisher Indole synthesis: Preparation of 2- phenylindole from phenylhydrazine. Enzymatic synthesis Enzymatic reduction :reduction of ethyl acetoacetate using Baker's yeast to yield enantiomeric excess of S(+) ethyl -3-hydroxybutanoate and determine its optical purity- \rightarrow Biosynthesis of ethanol from sucrose. Synthesis using microwave Alkylation of diethyl malonate with benzyl chloride- \rightarrow . Synthesis using phase transfer catalyst. Alkylation of diethyl malonate or ethyl acetoacetate with an alkyl halide.

Spectroscopy

Identification of organic compounds by the analysis of their spectral data (UV, IR, PMR, CMR & MS)

Spectrophotometric (UV/VIS) Estimations/ isolation of the following (any one compound)

Spectroscopic estimation

1. Amino acids
2. Proteins
3. Carbohydrates
4. Ascorbic acid
5. Aspirin
6. Caffeine

Isolation :

1. Casein from milk
2. Lycopene from tomatoes
3. Piperine from black pepper
4. Caffeine from tea leaves
5. Lactose from milk

Acero

Dr

Dr

Dr. Anila Dutt Rawar
Nasla

Dr

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(2023-24)

Class & Semester : M.Sc. IV
Subject : Chemistry
Paper : Chemistry Practical
Title : Physical Chemistry

Max Marks : 34

Time: 6-8hours

Experiment No.1	13
Experiment No.2	12
Record	04
Viva Voce	05
Total	34

Spectroscopy

- i. Determination of pK_a of an indicator (e.g. methyl red) in (a) aqueous and (b) micellar media.
- ii. Determination of stoichiometry and stability constant of Ferric isothiocyanate complex ion in solution.
- iii. Determination of rate constant of alkaline bleaching of Malachite green and effect of ionic strength on the rate of reaction.

Chemical Kinetics

- i. Determination of rate constant and formation constant of an intermediate complex in the reaction of Ce(IV) and Hypophosphorous acid at ambient temperature.
- ii. Determination of energy and enthalpy of activation in the reaction of $KMnO_4$ and benzyl alcohol in acid medium.
- iii. Determination of energy of activation and entropy of activation from a single kinetic run.
- iv. Kinetics of an enzyme catalyzed reaction.

Thermodynamics

- i. Determination of partial molar volume of solute (e.g. KCl) and solvent in a binary mixture.
- ii. Determination of the temperature dependence of the solubility of a compound in two solvents having similar intramolecular interactions (benzoic acid in water and in DMSO water mixture and calculate the partial molar heat of solution.)

Polarography

- i. Identification and estimation of metal ions such as Cd^{+2} , Pb^{+2} , Zn^{+2} , and Ni^{+2} etc. polarographically.
- ii. Study of a metal ligand complex polarographically (using Lingane's Method).

The students are expected to visit industries to gain practical knowledge, if possible.

Mune

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[Signature]

Mita Subhakar
Nawla
Anu